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paperback. Book Condition: New. Ship out in 2 business day, And Fast shipping, Free Tracking number will be provided after the shipment. Pages Number: 336 Publisher: Scientific Pub. Date :2010-10-01 version 1. This book is the physical chemistry of metallurgical processes thermodynamic part. A total of 12 chapters. Chapter 1 to 10. from the metallurgical point of view of thermodynamics. to discuss oxidation. reduction. melting matte blowing. chlorination. chemical transport. slag and molten steel reaction (desulfurization and dephosphorization). solidification of liquid steel (including oxygen. nuclei growth. to gas and segregation) and other physical and chemical principles of metallurgical processes. Implement the use of free energy book. extensive use of the isothermal equation and activity. analyze and solve the direction of metallurgical reaction equilibrium problems. Chapter 11. stating the thermodynamic parameters of the state diagram of the application; Chapter 12. summarizes the free energy calculation method. in addition to the general metallurgical reaction of the free energy calculation of the summary. but on the phase transition free energy. dissolved free energy and various standard state and activity of reference state calculations are discussed. with a view with a solution (that is. melt liquid metal metallurgy. slag. molten matte. etc.) to participate in the metallurgical reaction to a certain degree of quantitative analysis. Book for metallurgical engineering and related professional and technical personnel. research workers and tertiary students reference reading; also specialized courses as part of metallurgical physical chemistry textbooks. Contents: Preface Chapter symbol table book uses the basic laws of oxidation-reduction oxidation-reduction law 1-1 of the necessity and importance of 1-2 redox contradictory unity of opposites 1-3 oxide standard free energy F. And temperature T diagram and the use of 1-4 min extract with O- 1-5 direct reduction and indirect reduction 1-6 dissolved in liquid iron in the elements of the oxidation...



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